Deacon's Challenge No. 55 Answer

The upper limit of the reference range for mercury excretion in urine in occupationally exposed workers is given as: $10~\mu g~Hg/g$ creatinine.

Express this as nmol Hg/mmol creatinine.

(Atomic weight mercury 200.6, molecular weight creatinine 113.1).

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The general expression relating weights in mass units to SI units is:

It is simplest to consider concentrations of mercury and creatinine in turn.

Mercury is expressed as μg and we wish to convert it to nmol. Division of mercury in μg by its atomic weight (200.6) will give mercury in μmol . Since there are 1000 nmol in each μmol , this value must be multiplied by 1000.

Therefore for mercury:

$$\begin{array}{lll} \text{Mercury (nmol)} & = & \underline{\text{Mercury (\mu g)}} & x & 1000 \\ & & 200.6 & \end{array}$$

Creatinine is expressed as g/L and we wish to convert it to mmol. Division of creatinine in g by its molecular weight (113.1) will give creatinine in mol. Since there are 1000 mmol in each mol, this value must be multiplied by 1000.

Therefore for creatinine:

$$\begin{array}{rcl} \text{Creatinine (mmol)} & = & \underline{\text{Creatinine (g)}} & x & 1000 \\ & & 113.1 \end{array}$$

12 • ACB News Issue 510 • October 2005

Questions MRCPath Short Questions MRCPath Short Questions

These two expressions are combined in order to convert the mercury: creatinine ratio from $\mu g/g$ to nmol/mmol:

Note that since creatinine appears in the denominator its molecular weight appears in the numerator (since mercury concentration is divided by creatinine concentration). The 1000s cancel and substituting the mercury creatinine ratio in $\mu g/g$ gives:

Mercury:creatinine (nmol/mmol) =

 $\frac{10 \times 113.1}{200.6}$ = 5.6 nmol/mmol (2 sig figs)

Question 56

Calculate the loading dose of digoxin (bioavailability 0.75, salt factor = 1) required to achieve an initial plasma concentration of 1.5 μ g/L in a 60 kg man (assume volume of distribution = 7 L/Kg):

- a) If the patient has never taken digoxin
- b) If the patient is currently on digoxin with a plasma concentration of 0.5 $\mu\text{g}/L$